



## Double and Triple Covalent Bonding – Examples



Covalent Bond: Bond between two non-metals.

Double Covalent Bond: Two non-metals bonding together by sharing 2 unpaired electrons per atom.

Triple Covalent Bond: Two non-metals bonding together by sharing 3 unpaired electrons per atom.

| Atoms in the compound            | Lewis Dot Diagram of the Final Compound   | "Stick" Diagram  |
|----------------------------------|---|--|
| $O_2$<br>2 Oxygen Atoms          | $\begin{array}{c} \cdot\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{O}::\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \end{array}$   | $\begin{array}{c} \cdot\ddot{O}::\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \end{array}$                                    |
| $N_2$<br>2 Nitrogen Atoms        | $\begin{array}{c} \cdot\ddot{N}\cdot \\ \cdot\ddot{N}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{N}:::\ddot{N}\cdot \\ \cdot\ddot{N}\cdot \end{array}$  | $\begin{array}{c} \cdot\ddot{N}:::\ddot{N}\cdot \\ \cdot\ddot{N}\cdot \end{array}$                                   |
| $CO_2$<br>1 Carbon +<br>2 Oxygen | $\begin{array}{c} \cdot\ddot{C}\cdot \\ \cdot\ddot{O}\cdot \quad \cdot\ddot{O}\cdot \end{array} \rightarrow \begin{array}{c} \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \quad \cdot\ddot{O}\cdot \end{array}$ | $\begin{array}{c} \cdot\ddot{O}::\ddot{C}::\ddot{O}\cdot \\ \cdot\ddot{O}\cdot \quad \cdot\ddot{O}\cdot \end{array}$ |

# Dot Point Chemistry Answers Bonding

**Mary Millhollon, Richard H. Langley**



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