

# TYPES OF CHEMICAL REACTIONS

## DISPLACEMENT (SINGLE AND DOUBLE)

### INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.



Reaction: **SINGLE** DISPLACEMENT



Reaction: **DOUBLE** DISPLACEMENT

### QUESTIONS:

- 1)  $\_\_\text{Cl}_2 + \_\_\text{KI} \rightarrow \_\_\text{KCl} + \_\_\text{I}_2$  \_\_\_\_\_DISPLACEMENT
- 2)  $\_\_\text{K}_2\text{CO}_3 + \_\_\text{BaCl}_2 \rightarrow \_\_\text{KCl} + \_\_\text{BaCO}_3$  \_\_\_\_\_DISPLACEMENT
- 3)  $\_\_\text{Na} + \_\_\text{MgCl}_2 \rightarrow \_\_\text{NaCl} + \_\_\text{Mg}$  \_\_\_\_\_DISPLACEMENT
- 4)  $\_\_\text{Al} + \_\_\text{CuCl}_2 \rightarrow \_\_\text{AlCl}_3 + \_\_\text{Cu}$  \_\_\_\_\_DISPLACEMENT
- 5)  $\_\_\text{Al} + \_\_\text{Pb(NO}_3)_2 \rightarrow \_\_\text{Al(NO}_3)_3 + \_\_\text{Pb}$  \_\_\_\_\_DISPLACEMENT
- 6)  $\_\_\text{AgNO}_3 + \_\_\text{CaBr}_2 \rightarrow \_\_\text{AgBr} + \_\_\text{Ca(NO}_3)_2$  \_\_\_\_\_DISPLACEMENT
- 7)  $\_\_\text{Al}_2(\text{SO}_4)_3 + \_\_\text{Ca}_3(\text{PO}_4)_2 \rightarrow \_\_\text{AlPO}_4 + \_\_\text{CaSO}_4$  \_\_\_\_\_DISPLACEMENT
- 8)  $\_\_\text{FeBr}_2 + \_\_\text{K}_2\text{CO}_3 \rightarrow \_\_\text{FeCO}_3 + \_\_\text{KBr}$  \_\_\_\_\_DISPLACEMENT
- 9)  $\_\_\text{Cu} + \_\_\text{Ag}_2\text{SO}_4 \rightarrow \_\_\text{Ag} + \_\_\text{CuSO}_4$  \_\_\_\_\_DISPLACEMENT
- 10)  $\_\_\text{CuS} + \_\_\text{KCl} \rightarrow \_\_\text{CuCl}_2 + \_\_\text{K}_2\text{S}$  \_\_\_\_\_DISPLACEMENT

# Double Displacement Reactions Answer Key

**Matthew Johll**



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*Basic Chemistry* Steven S. Zumdahl, 1999 For a full description see catalog entry for Zumdahl

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