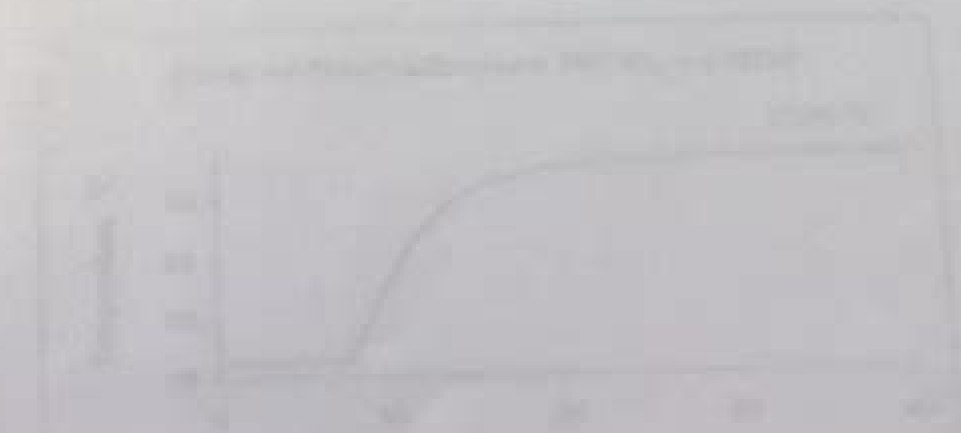


4. In accordance with Hess's Law, manipulate the following equations and their  $\Delta H_{\text{rxn}}$ 's to calculate  $\Delta H_{\text{rxn}}$  for the reaction  $\text{HClO}_4 + \text{NH}_3 \rightarrow \text{NH}_4\text{ClO}_4$ .

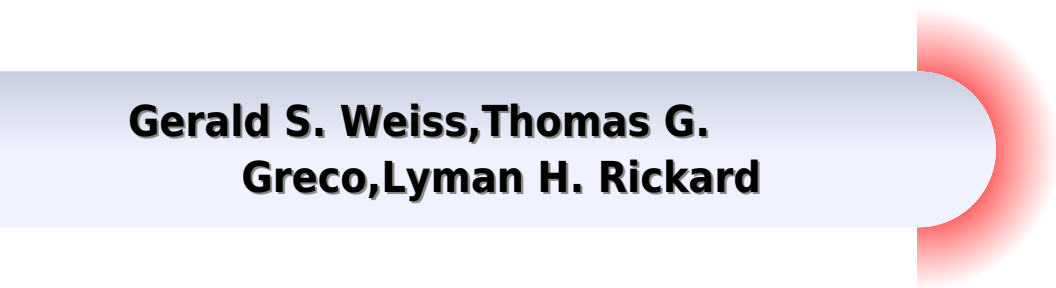


|                                                                                      |                                               |
|--------------------------------------------------------------------------------------|-----------------------------------------------|
| _____                                                                                | $\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$ |
| _____                                                                                | $\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$ |
| _____                                                                                | $\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$ |
| $\text{HClO}_{4(aq)} + \text{NH}_3_{(aq)} \rightarrow \text{NH}_4\text{ClO}_{4(aq)}$ | $\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$ |



# Experiments In General Chemistry Featuring Measurenet Answers

**Gerald S. Weiss, Thomas G.  
Greco, Lyman H. Rickard**



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